

Chapter 17 Thermochemistry Study Guide

17.1 – 17.2 Thermochemical Equations

Hansen Chemistry

- Make the following conversions:
 - 444 cal to joules = 1.86×10^3 J
 - 1.8 kJ to joules = 1.8×10^3 J
 - 0.45 kJ to calories = 1.1×10^2 cal
- Classify each of these processes as endothermic or exothermic:
 - condensing steam – **exo**
 - burning alcohol – **exo**
 - evaporating alcohol – **endo**
 - baking a potato – **endo**
- The specific heat capacity for silver is 0.24 J/g°C. Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(107.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

- It takes 585 J of energy to raise the temperature of 125.6 g Hg from 20.0°C to 53.5°C. Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

- A 46.2-g sample of copper is heated to 95.4°C and then placed in a calorimeter containing 75.0 g water at 19.6°C. The equilibrium temperature in the calorimeter is 21.8°C. Calculate the specific heat capacity of copper; assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

- A 15.0-g sample of nickel metal is heated to 100.0°C and dropped into 55.0 g of water, initially at 23.0°C. Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is 0.444 J/g°C.

$$\begin{aligned} -15.0(.444)(T_f - 100.0) &= 55.0(4.184)(T_f - 23.0) \\ -6.66T_f + 666 &= 230.12T_f - 5292.76 \\ 5958.76 &= 236.78 T_f \\ 25.2^\circ\text{C} &= T_f \end{aligned}$$

- Chloe was running bath water and realized it was too hot. If she has 20.0 L of water in the tub at 95°C and then adds 15.8 L of water at 75°C, what will the final temperature of the water be?

$$\begin{aligned} -20(T_f - 95) &= 15.8(T_f - 75) \\ -20T_f + 1900 &= 15.8 T_f - 1185 \\ 3085 &= 35.8 T_f \end{aligned}$$

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Thermochemistry Review Part 1. Definition Questions: you should be able to use or apply the following terms. Be able to . write. complete definitions for, energy temperature standard state . potential energy heat standard heat of formation. kinetic energy specific heat capacity exothermic. enthalpy endothermic Law of conservation of Energy . thermal Hess's Law isolated, closed, open system ...

Thermochemistry Review Worksheet

Answer to Problem 3) Thermo review: You are asked to design a flash drum to separate a 100 mole/min equimolar stream of hexane (1)...

Solved: Problem 3) Thermo Review: You Are Asked To Design ...

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Thermomix TM6 Review - Bake Play Smile

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Question: Problem 2) Thermo Review: On The Next Page Is A Pxy Diagram For A Mixture Of N-pentane And Methanol At A Temperature Of 422.6 K. Use The Information In The Figure To Answer Part A-f. A) Label The Regions Of The Diagram That Represent The Liquid, The Vapor And The VLE Phases. B) A Mixture Of 30 Mol% Pentane And 70 Mol% Methanol Is In The Vapor Phase ...

Problem 2) Thermo Review: On The Next Page Is A Px ...

Thermo - Problem 2.44 10 of 16 Review Constants Periodic Table Part A A 1.80 mol sample of an ideal gas for which $P_3=7.3$ bar and $T_3=28$ K is expanded adiabatically against an external pressure of 0.200 bar until the final pressure is 0.200 bar Calculate

the final temperature for $C_v=3R/2$.

Thermo - Problem 2.44 10 Of 16 Review Constants Pe ...

Thermo - Problem 2.53 12 of 16 A 1.10 mole sample of carbon dioxide, for which $C_{re-37.1}J-K-mollat$ 28 Kis expanded reversibly and adiabatically from a volume of 4.45 L and a temperature of 298 K to a final volume of 43.0 L Review Constants Periodic Table Part A Calculate the final temperature.

Solved: Thermo - Problem 2.53 12 Of 16 A 1.10 Mole Sample ...

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Lecture Notes on Thermodynamics - UMR 8550

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